

Date 20 MTWTFSS Difference between Floods of 2022 and super flood of 2010 Flood of 2022: The flood of 2022 were majorily the consequences of several reasons. The reasons were unusal unustral rainfall paterns, extreme weather conditions and Other Localissues. In some areas, climate change played major role. It intensified the frequency of flooding. But in Flood 2022, Their impacts and significance were majorly based upon the preparation and the infrastructure of the region. F100d 2010 : The flood of 2010, especially in Pakistan was called the super flood. Because of excessive moonsoon rown fall and river overfrow and inefficient drainage system caused the major outbreak of fload and local and Federal level authority were unable of handle these kinds of discuss is and in a result we with essed major flooding. Role of National Discuster Management Authority (NDMA) NDMA has the significant rov on manging the disauter risk and reponding to the emergencies. There ar multiple wayes, where they can marge - Preparation and planning, Resource management, coordination Jovernment between Local and provincia cuthoritics, Training and Awqueness gamppinges. Chawla Mote

Address all parts properly



그는 아님께서 말했다. 이 것은 것을 많은 것이 같은 것은 것을 것이 없는 것이 많이 많을까? 것을 하는 것을 수 있다. 것을 하는 것을 하는 것을 하는 것을 하는 것을 수 있다. 가지 않는 것을 하는 것을 하는 것을 수 있다. 가지 않는 것을 하는 것을 수 있다. 가지 않는 것을 하는 것을 수 있다. 가지 않는 것을 수 있다. 가지 않는 것을 하는 것을 수 있다. 가지 않는 것을 수 있다. 것을 수 있다. 가지 않는 것을 것을 수 있다. 가지 않는 것을 것을 것을 것을 수 있다. 가지 않는 것을 것을 것을 것을 것을 것을 수 있다. 것을		
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b) Differentiation 2		
(b) Differentiation Between stars and Planets		
1) Nature and company		
1) Nature and composition • Stars: "		
· Stars: Stars like the sun and others		
the massive caloriel have an energy. States are		
a emergine beat them have and they		
generate heat through nuclear fusion		
with helium and for atoms combine		
with helium and form multiple other atom with in the name accords and the		
the nano seconds and these reactions, releasing tremendous amount of energy.		
• <u>Plamets</u> : Plamets are composed of different		
material like gases, stare, rocks and ice etc.		
Planets revolve arount the stars - the		
bodies which orbit stars. and planets donot		
produce their own energy. They reflect the		
light from their stars.		
2) Size and Mass		
Stars are generally more massive than		
the plantets. Has of the sum. Their masses		
are typically more than the planets like		
sunmass is more than the earth and		
their diameter are also more larger like in		
thousand times them of the easth, on the		
other hand planets are smaller in size and		
also in mass and diameter. They range		
from 1000 to 10,000 from of kilumeter in		
diameter.		
Properly differentiate		





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3) life cycle :-	
Stars: Stars goes throug	h the lifecycle
and pass from various sta	ges like main
sequence, red giant and c	utimately end
with white draft or Black	hole and it depends
of their mass.	
Planets: Planets donot a	o through the phones
The the stores. They donot h	ave the lifecucio
similar to the stars. Their	change only necus
one to incernal and extern	nal forcer.
How a star becomes a b	lack hole
1) scellar Evolution: Those st	tous which are
more mansive like Do times m	ansive from the sum.
the process of them's end star	with stellar
evolution of becoming black	hole.
2) Supernova Explosion:- A	Nhen Nuclear fuel
of such massive stars exhau	sts then an a
explosion happen which we	called is super nova.
inclus the process where the	e core of the star
collapse under its own gr	Vilational force and
causing the explosion tou	Lards the outside of
3) Core collepse: After	the supernova
remaining core mass who	an reach at the
above certain level (openhim	omer limit tit will
furthmore collapse under i	ts own areats
growity and become bla	ck, thick dense
Steuch re.	
4) Formation of Blackhold	e - A blackhole.
is formed because fore	
dense and thick hat it	
Chanto	
Mare	
—	CSCa



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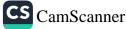
Date 20 MTWTFSS This is the point of infinite and this is swerounding event horizon from which even li can escape that why gravitational pull of Black from this hole is too strong (C) Part Why Atoms form chemical Bonds Atoms form the chemical bonds to become more stable by its electronic configuration. This stability can be achieved by Stable Valance shell: Atomic engage in chemical reaction bonds to fill there outer most el valance electrons to complete its electronic confeguration of like the Noble gares. These Nobles gares valence electrons are either 2 (duel) or 8 (octate). which is energetically more stable. Minimizing Energy:- Potential energy become Low during the process of bond 2) formation. By combing, these atoms loose their energy and leading towards the more stable form. Electron sharing and Transfer:-3) To achieve the full valance shells, atoms form the chemical bond either by Shaving (covalent Bond) or by complete transfer of electrons which is (Ionic



Representation?? 20 MTWTFSS tructue of water 1) Molecular Structure:-H20 is the formula for the water It- has the bend or v-shaped Structure. The bond angle between Hydrogen-oxygen-Hydrogen atomis approximately 104.50 2) Bonding - Covalent Bon 2s H20 (water) have 2 covalent bond with oxygen atom and each of the two with Hydrogenatoms. These involve the one pair Share of electrons between Hyd ogen & Oxygenatu 3) Polarity :- (Polar molecule H20 (water) is a polar molec len be cause Oxygen is the more electron gative then of the Hydrogen. This creater the partial pegative charge on the oxygen & partial positive charge on the Hydror in atom, leading to an overall dipole molecule. 4) Hydrogen Bonding: Master molecule con ain the strong Hydrogen he attraction b/ the & Bonding containing partial negative positive charge gen Hydrogen ato is of one mo ease of the partial negative of Jgen atom of another molecule. The Hydrogen bonding is responsible for High beiling Point heat capaç hawla

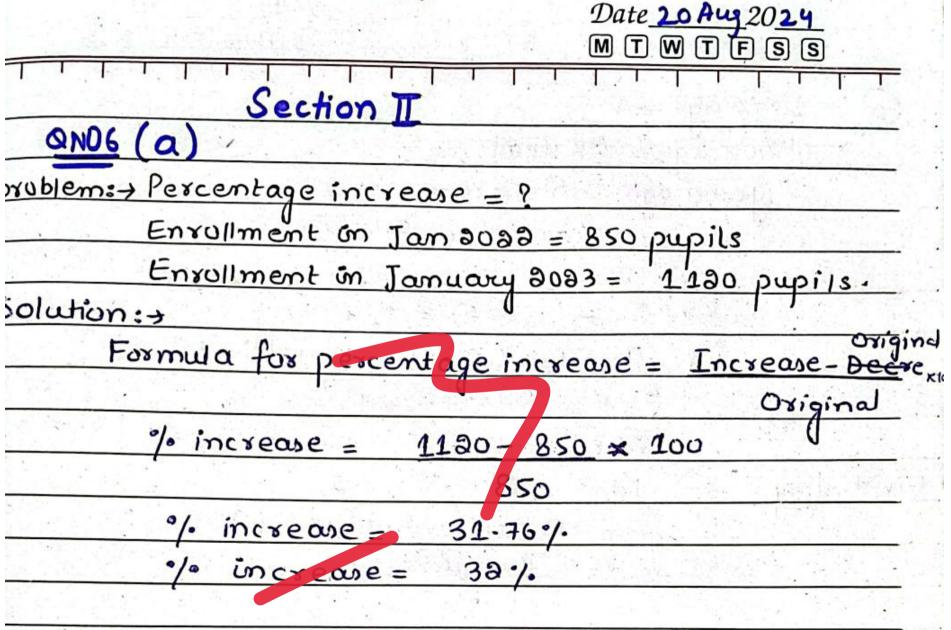


Date 20 MTWTFSS (0) is Conductors and casu Material that allow the free flow o electric current and heat due to free moving electrons Example: > copper is widely usedin electric wiring due to its electrical conductivity. in Semiconductors :- Material with electrical conductivity between insulators and conductors. The conductivity campe damaged or altered by adding impurities or environmental condition Example: > Silicon - used as computerchips. VID Metals :- Good conductor of Heat and electricity and shining in apperance Ex: > Aluminium -)' use in ackagin iv) Plastes: - materials made from polymers They are insulating & non conductive. Ex - Poly thylene -> Plastic ba

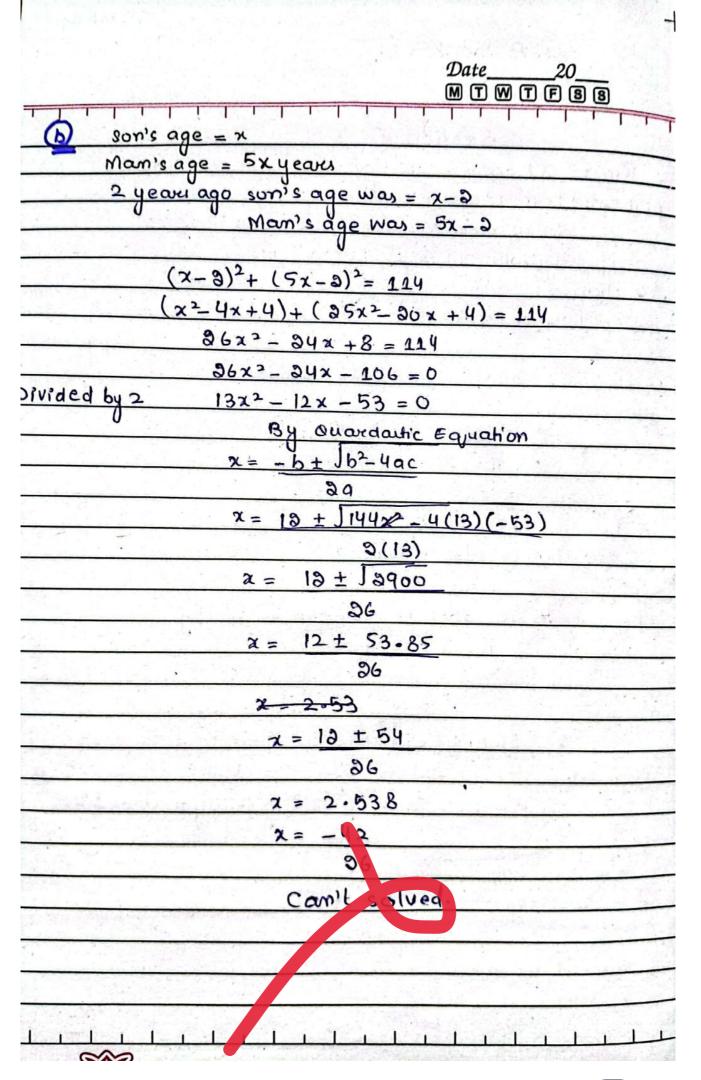


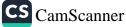
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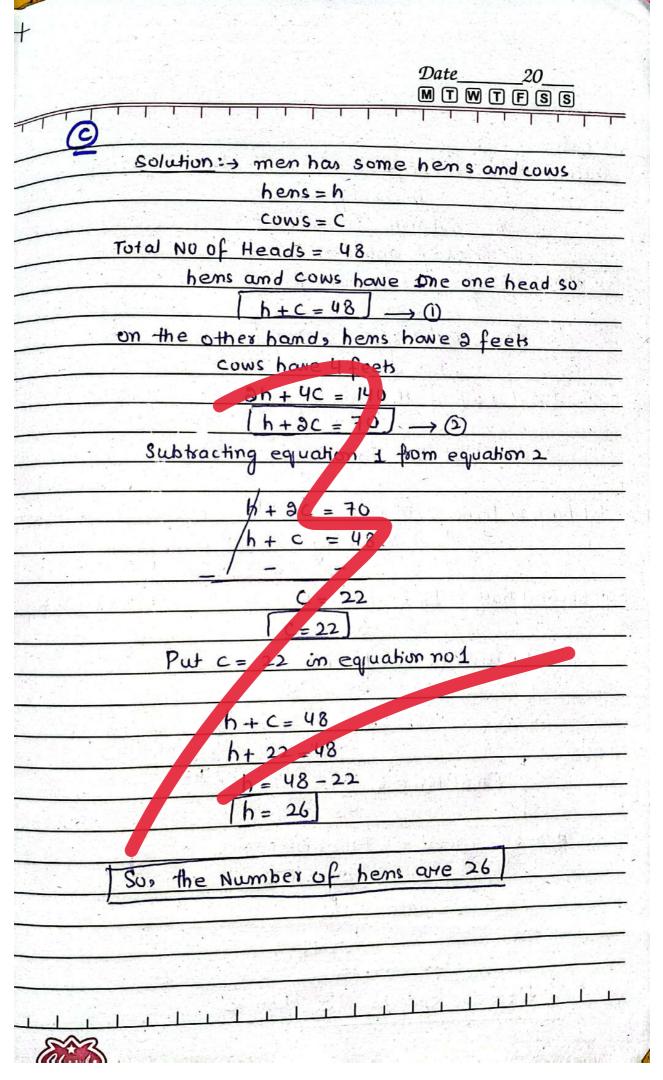
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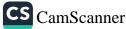












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